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PRODUCTION OF LIQUID SUSPENDED PHOSPHORUS NITERETRE

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Annotation

The main of chemical composition of liquid suspended of phosphoric ammonium nitrate fertilizers which obtained from central Kyzyl-kum's phosphorite with nitric acid has been determined. Optimal parameters of reception these fertilizers have been found and rheological properties of liquid suspended combined fertilizers have been studied.

Keywords. Porridge with phosphorite flour, nitric acid, calcium nitrate, ammonium nitrate, nitrocalcium phosphate

Currently, an urgent problem is the creation of complex liquid suspended fertilizers and fertilizer preparations that help increase the viability of plants to adverse environmental conditions and plant resistance to various diseases, as well as increase productivity.

It is known that liquid suspended fertilizers (LSF) are more promising and economical. They have a number of advantages over solid ones and fully meet the requirements of agrochemical science for the organization of highly efficient cultivation of crops.

In order to develop the theoretical foundations and technology for obtaining liquid suspended phosphorus-containing nitrate, we conducted research on the decomposition of unenriched phosphorite flour of the Central Kyzylkum at an incomplete rate of nitric acid in the presence of ammonium nitrate.

For the synthesis of new types of liquid complex suspended fertilizers, unenriched phosphorite flour of the composition was used, (wt.%): P₂O₅ - 17.55; CaO - 43.68; CO₂ - 14.83; MgO - 1.68; R₂O₃ - 2.47; SO₃ - 2.01; F - 2.17; H₂O - 1.19; but. - 3.80, 58.50% nitric acid of JSC "Maxam-Chirchik" and 64.16% ammonium nitrate



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solution obtained by neutralization of nitric acid with gaseous ammonia. The norm of nitric acid - 20-70% of the stoichiometry was calculated for the decomposition of phosphate and carbonate minerals of phosphorite flour to the formation of monocalcium phosphate and calcium nitrate.

Nitric acid decomposition of phosphate raw materials in a laboratory plant was carried out according to the developed intensive method. Phosphate with intensive stirring was treated with the calculated amount of nitric acid for 5-10 minutes. The calculated amount of ammonium nitrate solution was introduced into the resulting mass and stirred until a homogeneous suspension was obtained. Next, samples of liquid suspended complex fertilizers were analyzed for the content of basic nutrients according to well-known methods [1-20].

The study showed that the process of interaction of phosphorite flour with nitric acid proceeds intensively, practically without foaming, since it occurs, unlike classical methods, in the "solid-phase mode" and the resulting bubbles are destroyed in the phase separation. The structural features of granular phosphorites and the existence of three forms of carbonates in their composition [21-45] characterize the intense reactivity of these phosphorites.

During the nitric acid processing of high-carbonate phosphorites, the released carbon dioxide accelerates the decomposition of the phosphate mineral - fluorocarbonate apatite. The process is exothermic, the temperature of the reaction medium, depending on the norm of the acid, rises to 30-40°C. It has been established that with an increase in the rate of nitric acid, the degree of decomposition of phosphate rock increases.

Table 1 Chemical composition of liquid suspended phosphorus-containing nitrates, %

Ratio N:P ₂ O ₅	N _{general} .	P ₂ O ₅		CaO	
		general	water	general	water
Norm HNO ₃ , 20%					
1:0,5	13,25	6,63	-	16,50	2,67
1:0,6	12,33	7,39	-	18,18	2,95
1:0,7	11,39	7,97	-	19,85	3,22
1:0,85	10,40	8,84	-	21,64	3,50
1:1	9,39	9,39	-	23,41	3,76
Norm HNO ₃ , 30%					
1:0,5	13,05	6,52	-	16,24	3,92
1:0,6	12,12	7,27	-	17,87	4,32
1:0,7	11,18	7,82	-	19,48	4,70

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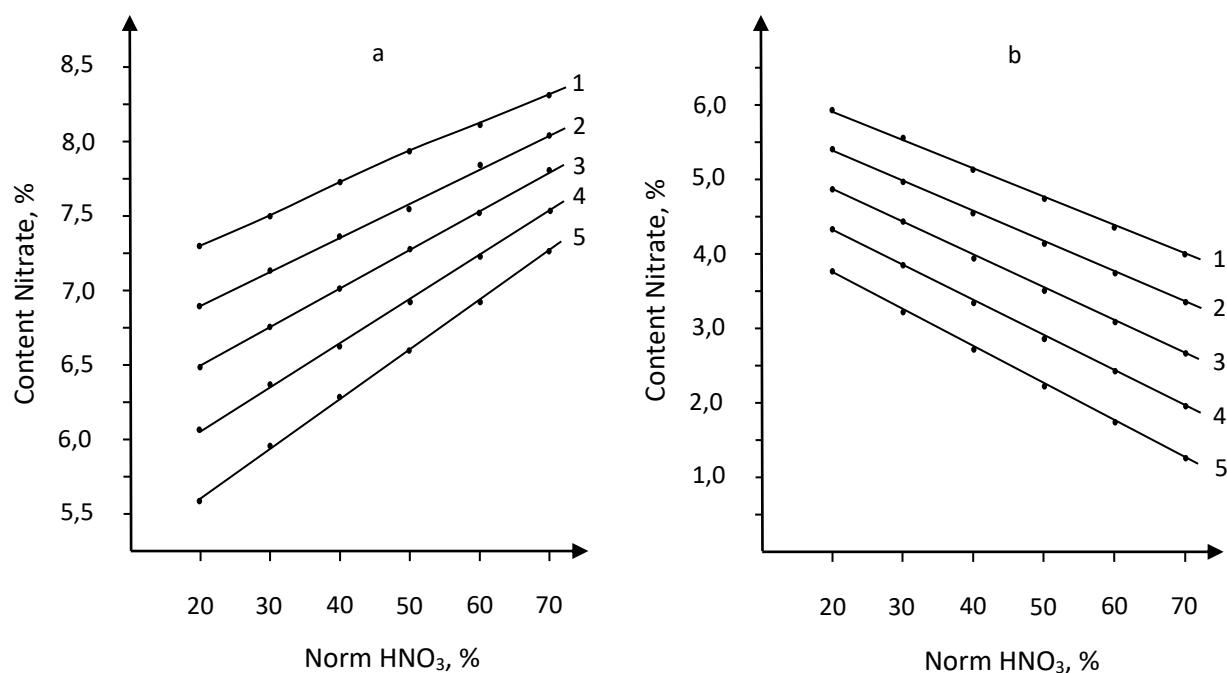
1:0,85	10,21	8,67	-	21,20	5,11
1:1	9,20	9,20	-	22,90	5,50
Norm HNO ₃ , 40%					
1:0,5	12,85	6,42	1,15	15,99	5,16
1:0,6	11,92	7,15	1,28	17,56	5,67
1:0,7	10,97	7,68	1,38	19,12	6,16
1:0,85	9,99	8,49	1,54	20,77	6,69
1:1	9,00	9,00	1,63	22,41	7,20
Norm HNO ₃ , 50%					
1:0,5	12,65	6,32	1,64	15,75	6,32
1:0,6	11,72	7,03	1,82	17,27	6,95
1:0,7	10,77	7,54	1,96	18,77	7,56
1:0,85	9,79	8,32	2,17	20,36	8,19
1:1	8,81	8,81	2,31	21,93	8,80
Norm HNO ₃ , 60%					
1:0,5	12,51	6,25	2,09	15,57	7,52
1:0,6	11,58	6,94	2,32	17,05	8,23
1:0,7	10,63	7,44	2,50	18,52	8,92
1:0,85	9,66	8,21	2,77	20,06	9,67
1:1	8,67	8,67	2,94	21,59	10,40
Norm HNO ₃ , 70%					
1:0,5	12,37	6,18	2,32	15,39	8,72
1:0,6	11,44	6,86	2,58	16,84	9,51
1:0,7	10,49	7,34	2,77	18,27	10,28
1:0,85	9,52	8,09	3,07	19,77	11,15
1:1	8,53	8,53	3,24	21,25	12,00

The main components of the products of nitric acid decomposition are tetrahydrate calcium nitrate, calcium phosphates, and activated phosphorite. The content of calcium nitrate increases with an increase in the rate of nitric acid.

To obtain a liquid suspended phosphorus-containing nitrate with different ratios of nutrients, the influence of the ammonium nitrate norm on the chemical composition and rheological properties of the product was studied.

The results of laboratory experiments are shown in Table 1. From the table data, it can be seen that at a ratio of N: P = 1: 0.5 and a norm of HNO₃ of 20%, the resulting fertilizer contains 13.25% N_{tot.}, 6.63% P_{2O5tot.}, 16, 50% Ca_{Tot.} and 3.87% CaO_{conv.}. With an increase in the N:P ratio to 1:1, the content of P_{2O5tot.} and CaO_{usv.} increases by 1.42 times and vice versa N_{tot.} decreases 1.41 times. In total, the sum of nutrient components (N+P_{2O5}+CaO_{usv.}) increases from 23.75 to 24.29%. A similar picture is observed at other norms of nitric acid. With an increase in the rate of HNO₃, the amount of nutrient components in the resulting liquid

suspended fertilizers increases. For example, at a ratio of N:P=1:1, with an increase in the norm of HNO₃ from 20 to 70%, the amount of nutrients increases from 24.29 to 32.75%.



Graph.1. The content of nitrate (a) and ammonia (b) nitrogen depending on the norm of HNO₃ at N:P 1–1:0,5; 2–1:0,6; 3–1:0,7; 4–1:0,85 и 5–1:1.

Figure 1 shows the change in the content of ammonia and nitrate nitrogen in the obtained suspended fertilizers depending on the rate of nitric acid at various N:P ratios. It can be seen from the figure that with an increase in the norm of nitric acid, the content of nitrate nitrogen increases, and ammonia nitrogen decreases. With an increase in the N:P ratio from 1:0.5 to 1:1, both forms of nitrogen decrease, since a large amount of phosphorite is introduced into the system.

Figure 2 shows the change in the assimilable forms of phosphorus and calcium in the obtained liquid suspended phosphorus-containing nitrates depending on the rate of nitric acid at various N:P ratios. It can be seen from the figure that with an increase in the rate of nitric acid, the content of assimilable forms of phosphorus and calcium increases, since with an increase in the rate of acid, phosphorite

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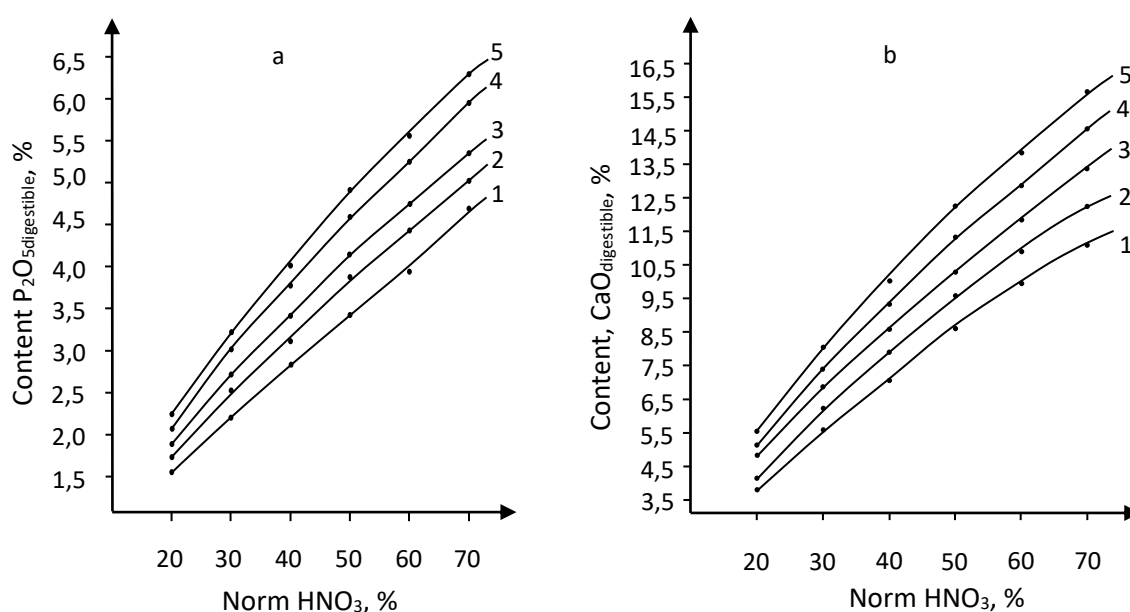
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decomposes more and dicalcium-, monocalcium phosphates and calcium nitrate are formed in quantity.



Graph.2. Content P₂O₅св. (a), CaOdigestible. (b) depending on the norm HNO₃ при N:P 1–1:0,5; 2–1:0,6; 3–1:0,7; 4–1:0,85 и 5–1:1.

An increase in the N:P ratio contributes to a decrease in digestible forms. In this case, the content of phosphorite increases in relation to nitric acid, and therefore the amount of dicalcium-, monocalcium-phosphates and calcium nitrate in the obtained fertilizers decreases.

Table 2 gives the salt composition of liquid fertilizers. From the tabular data it can be seen that with an increase in the N:P ratio, the salt composition of liquid suspended phosphorus-containing nitrate changes significantly. For example, at a ratio of N:P = 1:0.5 and a norm of HNO₃ of 20%, the composition of suspended liquids looks like this: activated phosphorite -30.22%, mono- and dicalcium phosphate -2.67%, calcium nitrate -7.75% and ammonium nitrate-34.10%. With an increase in the N:P ratio from 1:0.5 to 1:1 in the same norms of HNO₃ in the resulting liquid suspended fertilizers, the content of the main components increases, except for ammonium nitrate.

For example, at a rate of HNO₃ of 50%, with an increase in the N:P ratio from 1:0.5 to 1:1, the content of activated phosphorite, mono- and dicalcium phosphate and calcium nitrate increases by 1.39 times each, the content of

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ammonium nitrate decreases by 2.15 times. In all cases, the amount of the sum of nutrients and salts satisfies the requirements for liquid fertilizer in terms of nutrient concentration. In addition, when using such fertilizers, the main requirement is the saturated vapor pressure, the crystallization temperature of the solution and the dimension of the suspension, as well as the pH of the medium, since they predetermine the conditions for obtaining, storing, transporting and applying them to the soil. Despite this, such fertilizer can be applied not only through the roots, but through the leaves.

Table 1

Salt composition of liquid suspended phosphorus-containing nitrates

Ratio N:P	activated phosphorite	Mono-, dicalcium phosphate	Ca(NO ₃) ₂	NH ₄ NO ₃	H ₂ O
Norm HNO ₃ , 20%					
1:0,5	30,22	2,67	7,75	34,10	23,73
1:0,6	33,29	2,95	8,54	31,00	22,50
1:0,7	36,35	3,22	9,32	28,00	21,25
1:0,85	39,62	3,52	10,16	24,76	19,96
1:1	42,88	3,81	10,99	21,51	18,66
Norm HNO ₃ , 30%					
1:0,5	26,03	3,79	11,44	31,71	24,62
1:0,6	28,63	4,18	12,59	28,49	23,51
1:0,7	31,22	4,56	13,72	25,25	22,39
1:0,85	33,97	4,92	14,93	21,84	21,22
1:1	36,70	5,26	16,12	18,42	20,03
Norm HNO ₃ , 40%					
1:0,5	21,97	4,89	15,02	29,39	25,49
1:0,6	24,12	5,38	16,49	26,00	24,49
1:0,7	26,26	5,85	17,95	22,6	23,47
1:0,85	28,53	6,36	19,50	19,04	22,41
1:1	30,78	6,85	21,04	15,46	21,34
Norm HNO ₃ , 50%					
1:0,5	18,03	6,03	18,49	27,14	26,33
1:0,6	19,77	6,62	20,27	23,60	25,43
1:0,7	21,49	7,19	22,04	20,04	24,51
1:0,85	23,30	7,80	23,90	16,34	23,56
1:1	25,10	8,40	25,75	12,62	22,60
Norm HNO ₃ , 60%					
1:0,5	14,26	7,15	21,93	25,05	27,24
1:0,6	15,62	7,84	24,02	21,35	26,44
1:0,7	16,96	8,51	26,09	17,64	25,62



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1:0,85	18,37	9,22	28,26	13,80	22,31
1:1	19,77	9,92	30,41	9,94	18,98
Norm HNO ₃ , 70%					
1:0,5	10,49	8,27	25,37	22,96	28,15
1:0,6	11,47	9,06	27,76	19,11	27,45
1:0,7	12,43	9,83	30,14	15,24	26,73
1:0,85	13,44	10,64	32,61	11,26	21,05
1:1	14,44	11,44	35,07	7,26	15,36

In addition, the use of such types of fertilizers through the leaves of plants helps to increase the yield, resistance of plants to various diseases and bad weather. Therefore, it is necessary to study some of the physicochemical properties of these fertilizers in order to find the optimal parameters for obtaining such fertilizers that meet the above requirements.

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